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Determine the molecular weight and stoichiometric mole and mass air-fuel ratios for the Oklahoma gas mole composition given
in Table 3.4. Solution. Equation (3.2), $\text{CH}_4 + 2\text{O}_2 + 2(3.76)\text{N}_2\text{CO}_2 + 2\text{H}_2\text{O} + 2(3.76)\text{N}_2$ (3.2) shows that there are $2 + 2(3.76)$
 $= 9.52$ moles of air required for complete combustion of each mole of methane.

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