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$\ln \left(\frac{P}{P_{\text{vap}}} \right) = \frac{H_{\text{vap}}}{RT} - \frac{H_{\text{vap}}}{RT_{\text{K}}}$

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$N = \frac{P V}{R T} = \frac{18 \text{ mol} \times 8.314$

$\times 10^{-5} \text{ (bar} \cdot \text{m}^3$

$\text{mol} \cdot \text{K}^{-1})}{T} = \frac{P V}{R T} = \frac{1 \times$

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$10 - 3 - (10 - x) 106$

volume taken up by
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vap we have $x 18 \times$

$8.314 \times 10 - 5 T F$

$5432.8 \times l .$

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1.2 An approximate solution can be found if we combine Equations 1.4 and 1.5: $1.2 \text{ mV} \approx kT$ per molecular. Assume the temperature is 22 C.

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