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## Chapter 16

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*Homogeneous and Heterogeneous Mixtures Examples, Classification of Matter, Chemistry*  
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Pure Substances and Mixtures, Elements \u0026amp; Compounds, Classification of Matter,  
Chemistry Examples,

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3.2.1/3.2.2 Describe the differences between elements, compounds and mixtures. ~~States of~~  
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Science Experiments - Solid to Liquid, Liquid to Gas, Gas to Liquid Changes of States of  
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Solutions Manual Chemistry: Matter and Change • Chapter 9 145 CHAPTER 9 SOLUTIONS MANUAL 35. Aqueous solutions of potassium iodide and silver nitrate are mixed, forming the precipitate silver iodide. chemical equation:  $KI(aq) + AgNO_3(aq) \rightarrow KNO_3(aq) + AgI(s)$  complete ionic equation:  $K^+(aq) + I^-(aq) + Ag^+(aq) + NO_3^-(aq) \rightarrow K^+(aq) + NO_3^-(aq) + AgI(s)$

## **Chemical Reactions**

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Solution. Physical: boiling and melting are physical changes. When water boils, no bonds are broken or formed. The change could be written:  $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{g})$  Chemical: The dark grey nail changes color to form an orange flaky substance (the rust); this must be a chemical change.

## 3.6: Changes in Matter: Physical and Chemical Changes ...

240 Chemistry: Matter and Change • Chapter 12 Solutions Manual CHAPTER 12 SOLUTIONS MANUAL 30. Compare and contrast sublimation and evaporation. In both processes, the substances become a vapor. During sublimation, the substance goes from the solid phase directly to the vapor phase. During evaporation, particles in a liquid gain enough

## States of Matter

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