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11 Y 11 protons, 12 neutrons, 11 electrons 25 11 Y 11 protons, 14 neutrons, 11 electrons (2) iii) Average mass number = $[85 \ 100 \times 23] + [15 \ 100 \times 25] = 23.3$ (1) iv) They both contain the same number and arrangement of electrons. (1) v) They would have slightly different masses because 25 11 Y has two more neutrons than 23 11 Y, therefore 25 11 Y

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