

Chapter 9 Stoichiometry Worksheet

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~~CH Ideal Stoichiometric Calculations Chapter 9 2 Mr CChapter 9 Stoichiometry Worksheet~~
9.1. Objectives. •Definest Stoichiometry. •Describe the importance of the mole ratio in stoichiometric calculations. •Write a mole ratio relating two substances in a chemical equation. 2/26/2018 2. Stoich Definitions. •Composition stoichiometry deals with the mass relationships of elements in compounds. •Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

~~Chapter 9~~

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

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Stoichiometry Chapter 9, p. 275 - 294 Intro to Stoichiometry • Reaction Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction • Coefficients in a chemical reaction represent the mole ratios of each substance that react together • Example: 4Fe + 3O₂ → 2Fe₂O₃

~~Chapter 9 Stoichiometry Notes~~

9.3 Objectives Describe a method for determining which of two reactants is a limiting reactant. Calculate the amount in moles or mass in grams of a product, given the amounts in moles or masses in grams of two reactants, one of which is in excess.; Distinguish between theoretical yield, actual yield, and percentage yield.; Calculate percentage yield, given the actual yield and

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Chapter 9 Worksheet Stoichiometry 1. Consider the following equation: As₂O₃ + 2 Cl₂ + 5 H₂O → 2 H₃AsO₃ + 4 HCl a. How many moles of H₃AsO₃ can be produced from 3.37 mol of H₂O? b. How many grams of H₂O are needed to react with 0.789 mol of Cl₂? c. How many moles of As₂O₃ are needed to produce 62.7g of HCl? d.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g) N

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A

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