

## Chapter 11 Chemical Reactions Guided Reading Answers

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Chapter 11 Chemical Reactions113 SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329) This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction. Writing Chemical Equations (pages 321–323) 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.  $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$

### Chapter 11: Chemical Reactions

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Chapter 11 Chemical Reactions Guided Chapter 11 Chemical Reactions19 GUIDED PRACTICE PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction. Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and ...

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### Ch. 11: Chemical Reactions

Chapter 11 Chemical Reactions Study Guide - Quizlet Bookmark File PDF Chapter 11 Chemical Reactions Guided Reading AnswersChemical Reactions. In a chemical reaction, the reactants are written on the left and the products on the right. The arrow that separates them is called yield.

### [PDF] Chapter 11 Chemical Reactions

CHEMICAL REACTIONS. Chapter 11 Chemical Reactions 113. SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329). This section explains ..... GUIDED PRACTICE PROBLEM 17 (page 334). 17. Complete the equations for these single- replacement reactions in aqueous solution. Balance each equation. Write “ no reaction ” if a ...

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The initial step in solving a problem of this type must be to write the balanced chemical equation for the reaction. Inspection of Equation 11.4.1 shows that it is balanced as written, so we can proceed to the strategy outlined in Figure 11.4.1, adapting it as follows:

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